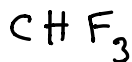
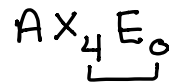
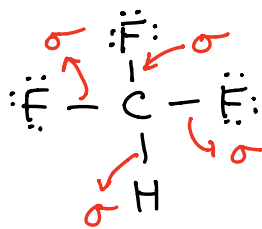
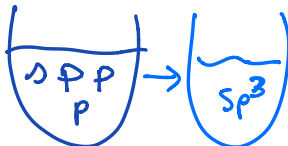
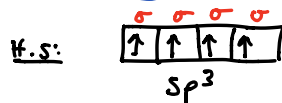
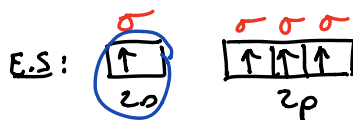
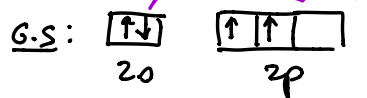


Hybridization Part II

Ex:



Carbon



$$4+0=4$$

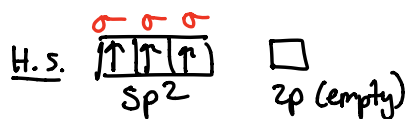
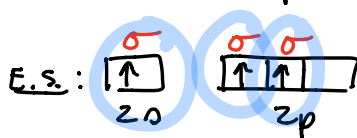
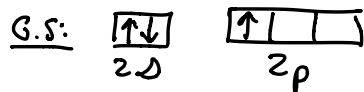
tells us how many hybrid orbitals we need

sp^2

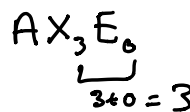
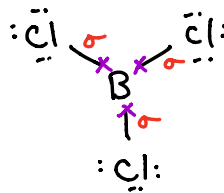
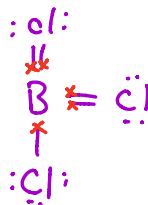
↳ mixing $s + p + p$ (sp^2)

Ex: BCl_3

Boron



Total v.e: $24 e^-$

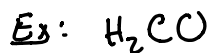


remaining e^- : $24 - 24 = 0$.

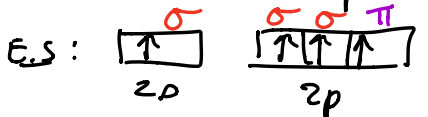
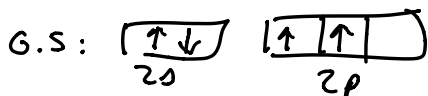
total used



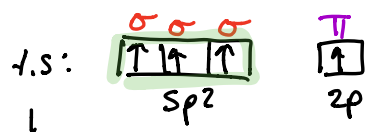
} sp^2 solution



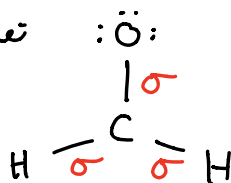
Carbon



↳ need to mix: $s + p + p$
 sp^2



Total: 12 val



remaining = $12 - 12 = 0$

AX_3

